

Molecular Compounds, Ionic Compounds, & Acids

NAME THE FOLLOWING COMPOUNDS:

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|---|---|
| 1. BaSO_3 barium sulfite | 14. AgNO_3 silver nitrate |
| 2. $(\text{NH}_4)_3\text{PO}_4$ ammonium phosphate | 15. As_2O_5 arsenic(V) oxide |
| 3. PBr_5 phosphorus pentabromide | 16. Fe_2O_3 iron(III) oxide |
| 4. MgSO_4 magnesium sulphate | 17. $\text{HClO}_{(aq)}$ hypochlorous acid |
| 5. CaO calcium oxide | 18. N_2O_3 dinitrogen trioxide |
| 6. $\text{H}_3\text{PO}_4_{(aq)}$ phosphoric acid | 19. $\text{HF}_{(aq)}$ hydrofluoric acid |
| 7. $\text{Na}_2\text{Cr}_2\text{O}_7$ sodium dichromate | 20. $\text{H}_2\text{C}_2\text{O}_4_{(aq)}$ oxalic acid |
| 8. MgO magnesium oxide | 21. NaHCO_3 sodium hydrogen carbonate |
| 9. SO_3 sulfur trioxide | 22. SiBr_4 silicon tetrabromide |
| 10. $\text{Cu}(\text{NO}_3)_2$ copper(II) nitrate | 23. CuCl_2 copper(II) chloride |
| 11. $\text{HI}_{(aq)}$ hydroiodic acid | 24. $\text{HNO}_2_{(aq)}$ nitrous acid |
| 12. N_2O dinitrogen monoxide | 25. SnO_2 tin(IV) oxide |
| 13. MnO manganese(II) oxide | 26. BaCrO_4 barium chromate |

WRITE FORMULAS FOR THE FOLLOWING COMPOUNDS:

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| 27. hydrobromic acid HBr | 40. diphosphorus pentoxide P_2O_5 |
| 28. chromium(III) carbonate $\text{Cr}_2(\text{CO}_3)_3$ | 41. sulfurous acid H_2SO_3 |
| 29. magnesium sulfide MgS | 42. lead(II) nitrate $\text{Pb}(\text{NO}_3)_2$ |
| 30. iodine trichloride ICl_3 | 43. dihydrogen monoxide H_2O |
| 31. lithium hydride LiH | 44. sodium oxalate $\text{Na}_2\text{C}_2\text{O}_4$ |
| 32. ammonium hydroxide NH_4OH | 45. perchloric acid $\text{HClO}_4_{(aq)}$ |
| 33. calcium chloride CaCl_2 | 46. chlorous acid $\text{HClO}_2_{(aq)}$ |
| 34. hydroselenic acid H_2Se [not possible] | 47. silicon dioxide SiO_2 |
| 35. iron(II) nitride Fe_3N_2 | 48. carbonic acid $\text{H}_2\text{CO}_3_{(aq)}$ |
| 36. aluminum hydroxide $\text{Al}(\text{OH})_3$ | 49. sodium chlorate NaClO_3 |
| 37. tin(II) fluoride SnF_2 | 50. xenon hexafluoride XeF_6 |
| 38. sulfur tetrachloride SCl_4 | 51. nickel(II) nitrate $\text{Ni}(\text{NO}_3)_2$ |
| 39. mercury(II) iodide HgI_2 | 52. potassium perchlorate KClO_4 |
| 53. stannic carbonite pentahydrate $\text{Sn}(\text{CO}_3)_2 \cdot 5\text{H}_2\text{O}$ | 54. lithium peroxide Li_2O_2 |
| 55. ferric hydrogen perphosphate $\text{Fe}_2(\text{HPO}_5)_3$ | 56. $\text{Cr}_2(\text{CO})_3 \cdot 3\text{H}_2\text{O}$ |
| 57. MnHPO_3 manganese(II) hydrogen phosphite | 58. MgCr_2O_4 chromium(III) hypocarbonate trihydrate
magnesium chromate |