

Ch. 8 + 9 Assignment ²⁵

$$10 \text{ mL} = 9.24 \text{ g}$$

1. a) $\frac{1.77 \text{ g}}{10 \text{ g}}$
b) $1.77 \times 9.24 \div 10 = \underline{1.64 \text{ g}}$
c) $1.77 \text{ g} / 10.8225 = \underline{16.35\%}$

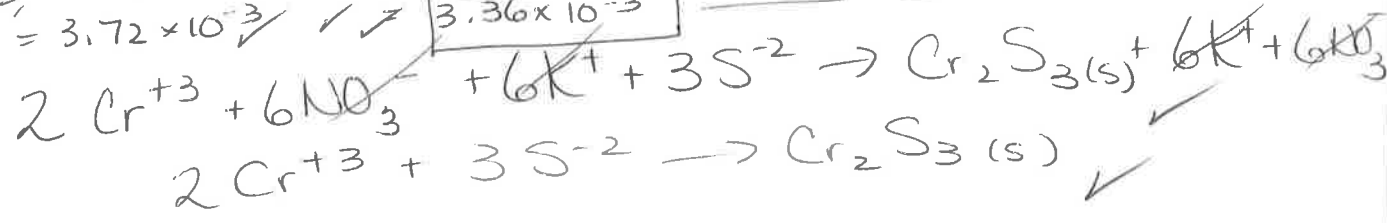
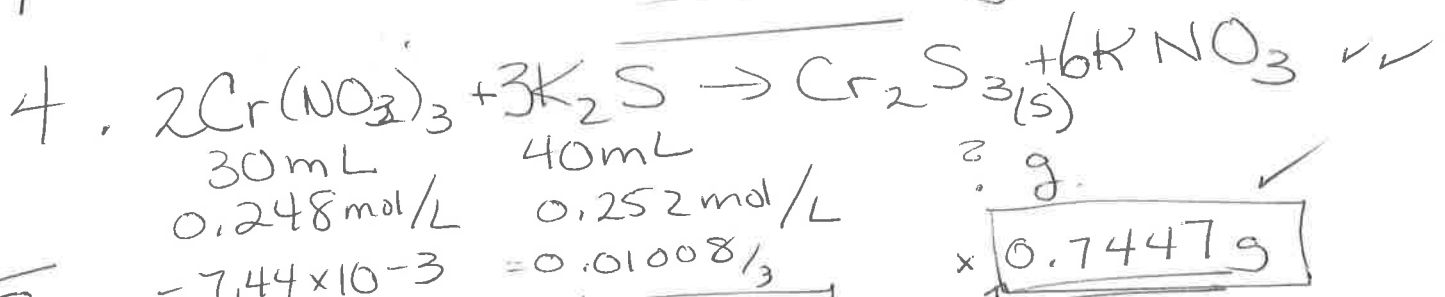
$$\frac{9.24 \text{ g}}{10 \text{ mL}} = \frac{10 \text{ g}}{x}$$

2. $C_1 V_1 = C_2 V_2$
 $500(0.555) = 0.255(V)$

$V = 1088 \text{ mL}$

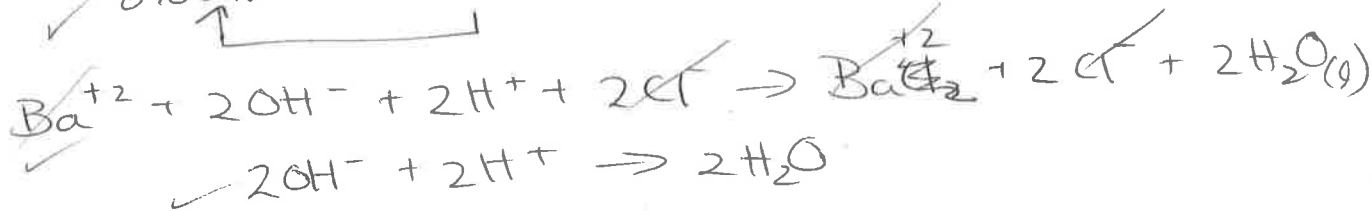
3. $\frac{0.24 \text{ mol}}{\text{L}} \times 0.05 \text{ L} \quad 2\% \times 150$
 $= 0.012 \text{ mol} \quad = 3 \text{ g}$
 $\times 74.55$
 $= 0.8946 \text{ g} +$

$= 3.8946 \text{ g KCl}$





\checkmark $\cancel{2.88} \text{ mol/L}$ 25 mL
 \checkmark $V = 0.01736 \text{ L}$ 4 mol/L
 \checkmark 0.05 mol $= 0.1 \text{ mol}$ ✓



~~c)~~ d) Neutralization, D D ✓

$(11) \rightarrow \dots$

$\dots = C_2V_2$